

Article

In Situ Effectiveness of Alkaline and Cementitious Amendments to Stabilize Oxidized Acid-Generating Tailings

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Abstract: This study investigates the effectiveness of alkaline and cementitious additives in the in situ stabilization of localized acid-generating tailings from a closed gold mine in Abitibi-Témiscamingue, Québec (Eagle/Telbel mine site). Five field cells (including one control) were constructed and equipped with mechanisms for collecting vertical water infiltration and surface runoff. The five cells included: (C1) Control cell; (C2) 5 wt % limestone amendment; (C3) 10 wt % limestone amendment; (C4) 5 wt % half ordinary Portland cement and half fly ash amendment; and (C5) 5 wt % ordinary Portland cement amendment. The control cell showed an acidic behavior ($\text{pH} < 4.5$) with variable concentrations of Fe, Al, Zn, and Cu. The amendments were used to neutralize the acidic leachates and decrease dissolved metal concentrations. Leachates from surface runoff samples of amended cells were less loaded with metals compared to samples of vertical infiltration. All amendment formulations increased the pH of the leachates from approximately 4 to circumneutral values. Furthermore, metal and metalloid concentrations were greatly limited, except for Cr and As for the carbonate-based amendments. Metal(-oid) stabilization was successfully achieved using the different amendment formulations, with the exception of C2, which still released As.

Keywords: acid mine drainage; alkaline amendments; cementitious amendments; kinetic testing; Joutel mine

1. Introduction

In Canada, environmental regulations require restoration and stabilization of tailings and effluents before the final closure of mine sites. The stabilization of acid-generating tailings and contaminants released from oxidized tailings can be achieved using stabilization/solidification (S/S) techniques. Such techniques consist of adding amendments, such as alkaline materials, cementitious materials, and/or industrial sub-products, to mine wastes [1–14]. Mining amendments can be divided into two categories, depending on the objective of the amendments: (i) Alkaline amendments, which are used for increasing the neutralization potential of acid-generating tailings [15–17], and (ii) cementitious amendments, which are used for the impermeabilization/solidification of acidic tailings, but can also increase the neutralization potential [18–23].

Alkaline amendments were successfully used to control and neutralize acid mine drainage under laboratory conditions during kinetic tests [1,8,11,24–27]. Limestone is one of the most commonly used materials in alkaline amendments. The dissolution of alkaline amendments under acidic conditions increases the alkalinity of leachates and reduces the mobility of contaminants through mechanisms

including: (i) Precipitation of low-solubility iron oxyhydroxides; (ii) co-precipitation in and adsorption on oxyhydroxides [11,28–31]; (iii) sulfide surface passivation, which inhibits oxygen diffusion to and from reactive sulfide cores [32–35]; and (iv) reduction of bacterial activity under circumneutral conditions [36–38]. Other materials that have successfully been used as alkaline amendments include red mud bauxite and cement kiln dust [11,12,39,40].

Cementitious amendments are used to solidify hazardous wastes and contaminated soils [7,20,41,42]. Cementitious amendments can reduce the contaminant mobility by the same mechanisms as alkaline amendments. However, the application of this technique allows for the solidification of reactive grains and ensures the physical trapping of contaminants (i.e., fixation). Moreover, the cementation of tailings leads to: (i) Physical encapsulation of mobile contaminants by increasing the cohesive properties of the mixture and (ii) improvements to the cohesion and the long-term impermeability of the tailings, which reduce available reactive surface area [4,7,20,22,23,42–45]. Various materials have successfully been used as cementitious additives for these purposes, although ordinary Portland cement (OPC) is the most common. In some cases, industrial sub-products, such as fly ash (FA), slag, lime, and cement kiln dust, have been used to partially replace OPC in cemented paste backfills [1,3,5,44,46–52].

In this study, limestone, fly ash, and ordinary Portland cement were tested at the former Eagle/Telbel mine site, which is now referred to as Joutel mine. The main objective of this study was to test the in situ effectiveness of these materials to stabilize acid-generating tailings. To accomplish this, five experimental cells were constructed and equipped with two types of water collectors, one for vertical infiltration and one for surface/subsurface runoff. This instrumentation allowed for time-series measurements of the chemical compositions of drainage waters, as well as calculations of the two main components of the water balance.

2. Materials and Methods

2.1. Joutel Mine Site

Joutel is a closed gold mine site located at the north of Abitibi–Témiscamingue (Québec, Canada). Operations at the mine took place between 1974 and 1994. The gold was associated with a sulfidic deposit that contained mainly pyrite, as well as traces of pyrrhotite, chalcopyrite, sphalerite, and galena. Gold was extracted using sulfide bulk flotation, followed by cyanidation. Ore treatment produced finely ground tailings that were deposited over 120 ha in a tailings storage facility (TSF) [53]. The TSF is divided into two zones: The northern zone, which is older and relatively elevated, and the southern zone, which is younger [54].

2.2. Amendments Formulation

Both alkaline and cementitious amendments were tested in in situ experimental cells. The alkaline amendments consisted of limestone. Two formulations were tested based on preliminary laboratory tests, i.e., 5 wt % and 10 wt % limestone. The amendment was mixed in situ with oxidized tailings and then deposited in the field cells. The limestone used in this study had a maximum particle size of 6.25 mm to ensure a mixture of fine and coarse particle sizes that would provide both short- and long-term reactivity. Calculations of the amount of limestone needed to neutralize the oxidized tailings were based on the neutralization potential (NP) and acid-generating potential (AP) of the limestone and the reactive tailings, as expressed in Equation (1):

$$\%R = 100 \times \frac{(\text{NP}_{\text{tailings}} - (f \times \text{AP}_{\text{tailings}}))}{(f \times \sum_{i=1}^n X_i \times \text{AP}_i) - (\sum_{i=1}^n X_i \times \text{NP}_i)} \quad (1)$$

where $\text{NP}_{\text{tailings}}$ is the neutralization potential of the mine tailings; $\text{AP}_{\text{tailings}}$ is the acidification potential of the mine tailings; f is the target NP/AP ratio; X_i is the proportion of each amendment material used (equals 1 if only one amendment is used); AP_i is the acidification potential of the amendment

material; and NP_i is the neutralization potential of the amendment material. For the studied tailings, the calculated ratio was ~8% for an NP/AP ratio of 1.

Cementitious amendments in this study consisted of OPC and FA. The two formulations used were 1/2 OPC + 1/2 FA 5 wt % and OPC 5 wt % at a dosage relative to the total dry weight of the tailings. The FA used in this study was taken from the Boralex thermal energy station (Senneterre, QC, Canada), which generates wood-residue thermal energy. The amendment formulations for the five cells are indicated in Table 1.

Table 1. Amendment formulations used in the field cells.

Cell	Formulation
C1	Reference (oxidized tailings)
C2	5 wt % limestone
C3	10 wt % limestone
C4	5 wt % (1/2 OPC + 1/2 FA)
C5	5 wt % OPC

2.3. Field Cell Construction

Field cells were constructed in an acidic area of the Joutel TSF. The cells, which were shaped as inverted truncated pyramids, were 4 m wide, 4 m long, and 0.3 m deep. Only oxidized tailings were used in the cells. The tailings were homogenized in situ using a mechanical loader to avoid, as much as possible, heterogeneities in the chemical and mineralogical properties of the materials. Amendments were mixed into the tailings using the bucket of a mechanical loader (Figure S1E–H). The cells were then excavated in the TSF and linear low-density polyethylene geomembranes were installed at the base and on the sides of each cell to control exfiltration [55] (Figure S1A,B,D,F). Two systems of water collection were installed at each cell. One system collected vertical infiltration, using 5-cm PVC pipes, and the other collected surface and subsurface runoff, using a combination of 5-cm PVC pipes and a gutter (Figure S1C). Each collection system was connected to an external reservoir. The cells using binders were covered with 20 cm of sand as a protection layer (Figure S1I). The drains for the collection systems were installed at a 2% slope. A schematic representation of the cells is shown in Figure S2. Each cell was also equipped with an EC5 volumetric water content (VWC) probe, which obtains water content by measuring the dielectric constant of the media through capacitance/frequency domain technology. Each probe was calibrated in the laboratory, using the same material present in the field. Calibration curves are presented in the supplementary materials (Figure S3). Sampling was not possible during winter; average daily temperatures are presented in the supplementary materials (Figure S4).

2.4. Physical, Chemical, and Mineralogical Analyses

Grain-size distributions of the studied samples were evaluated using a laser analyzer (Malvern Mastersizer, Malvern Panalytical, Canada). Bulk chemical compositions of the samples were determined by ICP-AES (Perkin Elmer Optima 3100 RL, USA), following an $HNO_3/Br_2/HF/HCl$ digestion. Major minerals in the initial oxidized tailings were analyzed by X-ray diffraction (XRD, Bruker Ltd, Canada; Bruker D8 Advance, with a detection limit and precision of approximately 1–5%, operating with a copper cathode, $K\alpha$ radiation). Results were interpreted using the DIFFRACT.EVA software (Version 3.1, Bruker, Milton, ON, Canada) and quantified using TOPAS (Version 4.2, Bruker, Milton, ON, Canada).

Leachates collected from the cells were analyzed for pH, Eh, and electrical conductivity (EC), using pH/Eh/EC meters, and their chemical compositions were analyzed by ICP-AES, following an addition of 2% HNO_3 . Iron–sulfur pH–Eh diagrams were calculated for the control cell and for the four cells with alkaline and cementitious amendments at 21 °C, using the Geochemist’s workbench (GWB) database (student edition, Version 12.0.1, Aqueous Solutions LLC, Richmond, VA, USA). Iron and

sulfur activities were calculated using Visual Minteq (Version 3.1, KTH, Stockholm, Sweden) based on average Fe and S concentrations. Then, the pH and Eh data from the different cells were projected onto the Fe–S pH–Eh diagrams.

After two years, a composite solid sample was collected from each cell. The chemical composition of these samples was analyzed using ICP-AES. The mineralogy of these samples was investigated using Quantitative Evaluation of Minerals by Scanning Electron Microscopy (QEMSCAN, FEI, Hillsboro, OR, USA). QEMSCAN is an automated mineralogy system that produces particle maps through the collection of rapidly acquired X-rays. The maps and corresponding data quantify modal mineralogy, texture, grain sizes, mineral liberation, and elemental department. In this study, carbonate and sulfide liberation were evaluated in order to quantify the coating of each phase and, ultimately, to determine the long-term geochemical behavior of the amended tailings (especially for the limestone amendments). Polished sections were prepared for each sample and analyzed using particle mineralogy analysis mode (PMA). Measurement resolutions varied from 2.5 μm to 6 μm , depending on the particle-size distribution of the sample. The number of analyzed particles was approximately 47391, 78089, 43077, 41851, and 42699 for samples from C1, C2, C3, C4, and C5, respectively.

3. Results

3.1. Chemical, Mineralogical, Static Test, and Physical Characterizations of Solid Samples

Table 2 shows the results of the chemical, mineralogical, and static test, and physical characterizations of the mine tailings, limestone, OPC, and fly ash. The grain-size distribution (GSD) of the mine tailings was finer compared to the limestone, OPC, and FA. The D_{90} , which corresponds to 90 wt % passing on the cumulative GSD curve, was approximately 30.5 μm for the mine tailings sample, 4500 μm for limestone, 46.7 μm for the OPC, and 1500 μm for the FA. Initial water content was about 15 wt % for the mine tailings sample, 6.1 wt % for the limestone sample, 0 wt % for the OPC, and 3.5 wt % for the fly ash. Mineralogical analyses showed that the mine tailings were highly oxidized. They were mainly composed of secondary minerals, such as goethite (22 wt %). Carbonates and sulfides constituted about 18 and 12 wt % of the tailings, respectively. The limestone sample was mostly composed of calcite (76 wt %) and dolomite (22 wt %). The fly ash sample contained about 15.5 wt % calcite, 3.5 wt % siderite, 5.3 wt % quartz, and 75 wt % of various silicate minerals. The reconciliation of the mineralogical composition of fly ash was based only on wt % of the crystalline phases. However, fly ash is known to contain amorphous alumina–silicate glass. Consequently, the sum of wt % of phases was less than 100 wt %. Chemical characterizations showed that the mine tailings sample was high in Fe (27%), S (4.40%), and Ca (3.7%), confirming the mineralogical composition. The limestone sample was composed mainly of Ca (34%) and Mg (2.50%). The OPC sample contained mostly Ca (49%), Al (2.75%), and Fe (2.25%). The fly ash sample contained mainly Ca (7.50%), Al (4.75%), and Fe (2.25%). Other elements were present as trace concentrations within the four samples and the complete chemical composition of these samples is presented in Table 2. NP values, which were determined using the Sobek method as modified by Bouzahzah et al. (2015) [56], were ~183 kg CaCO_3/t and 880 kg CaCO_3/t for the tailings and limestone, respectively. The AP of the tailings was ~195 kg CaCO_3/t .

Table 2. Physical, chemical, and static test, and mineralogical characterizations of the mine tailings, limestone, ordinary Portland cement, and fly ash samples. DL = detection limit.

Characterization	Parameter	Units	MT	Limestone	OPC	FA
Physical characteristics	D10		1.8	50	4.2	82
	D30	µm	4.7	250	11.3	180
	D90		30.4	4500	46.7	1500
	initial water content	wt %	15	6.1	dry	3.52
Chemical composition	Al		2.02	0.295	2.75	4.69
	Ca		3.72	33.82	49.07	7.58
	Mg		0.15	2.350	1.18	1.05
	Mn		0.42	0.031	0.06	0.42
	Na		1.07	0.125	0.16	1.96
	K		0.26	0.245	0.43	1.85
	Fe		27.24	0.483	2.23	2.24
	Si*		17.76		7.76	14.82
	Li	%	≤0.0005	≤0.0005	≤0.0005	0.002
	Pb		0.013	≤0.0005	≤0.0005	≤0.0005
	As		0.091	≤0.0005	0.005	≤0.0005
	Cr		0.0024	0.004	0.007	0.006
	Cu		≤0.001	≤0.0001	0.007	≤0.0001
	Zn		0.01	≤0.0055	0.05	0.07
	S (total)		6.80	0.93	1.74	0.43
S (sulfates)		0.27	-	-	-	
C		2.2	-	-	-	
Static tests	NP	kg CaCO ₃ /t	183	880	-	-
	AP	kg CaCO ₃ /t	195	-	-	-
Mineralogical composition	Quartz		27.84	2.07		5.32
	Calcite		7.82	75.76		15.51
	Dolomite		0.60	22.17		
	Muscovite		1.03			
	Siderite		11.41			3.65
	Orthoclase		0.99			8.34
	Biotite	wt %	0.23			7.52
	Albite		12.78		Not applicable	20.51
	Gypsum		1.45			
	Goethite		22.29			
	Pyrite		12.22			
	Labradorite		0.30			22.88
	Chlorite		0.74			
	Corundum		0.41			
Anhydrite					1.83	

Si* was analyzed using XRF-whole rock.

3.2. Field Cell Monitoring

3.2.1. Vertical Infiltration

Vertical infiltration leachates from the different cells showed different behaviors with respect to pH and EC (Figure 1A,B). Leachates from the reference cell were acidic, with pH values between 1.7 and 4.3. However, the leachates from the amended tailings were circumneutral, with pH values between 6.7 and 8.4 for C2, 7.1 and 8.1 for C3, 6.6 and 8.6 for C4, and 8.8 and 10.3 for C5 (Figure 1A). All amendment formulations successfully buffered the acid produced by the oxidized tailings. The acidic leachates react with the limestone and binders, which dissolve and neutralize acidity [7,11,17,26,31,57–59]. Eh values showed oxidizing conditions (Eh > 100 mV) in all five cells (Figure 1C). Electrical conductivity values (Figure 1B), which illustrate the chemical quality of the leachates, showed that leachates from C1 were more loaded in terms of dissolved ions. EC values for C1 ranged between 8 and 33 mS/cm. In contrast, EC values from C2, C3, C4, and C5 were between 5 and 20 mS/cm, 5 and 12 mS/cm, 1 and 7 mS/cm, and 2 and 5 mS/cm, respectively. The cementitious amendments showed the lowest EC values compared to the control cell and limestone-amended cells; hardening of tailings, due to

cementation processes, reduces the available surface area, which reduces the leaching rates of species from mineral surfaces [4,7,10,20,43,52,60,61].

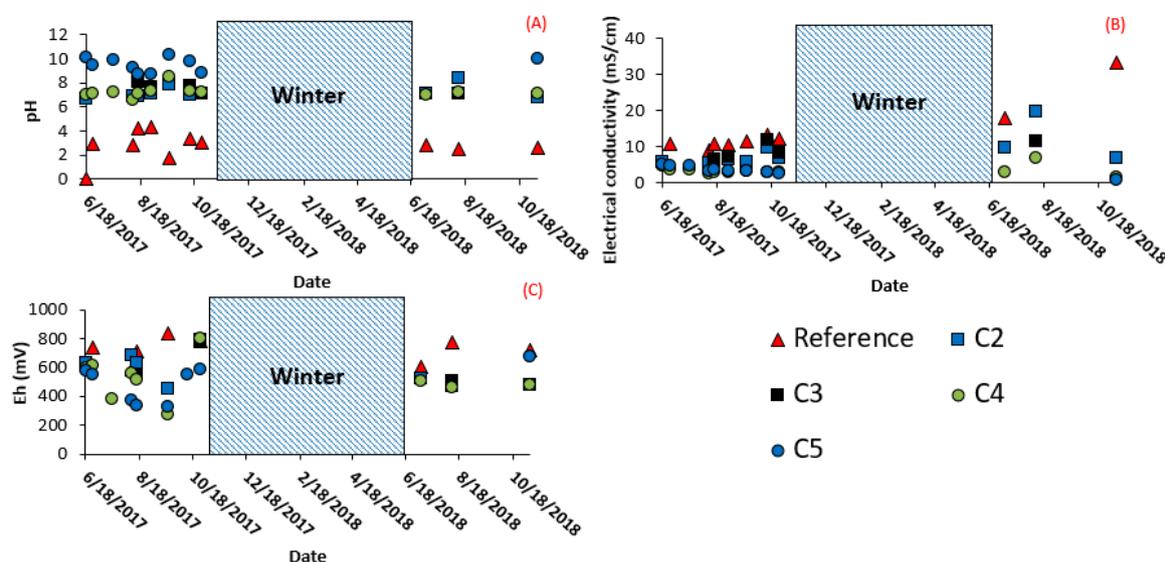


Figure 1. Evolution of (A) pH, (B) electrical conductivity, and (C) Eh for leachates from vertical infiltration.

Vertical infiltration leachates from the five cells were analyzed for concentrations of Ca, Mg, Mn, Al, Fe, S, Zn, Cu, As, and Cr to evaluate carbonate dissolution, sulfide oxidation, and metal(-oid) release rates (Figure 2). The end of the test was characterized by increased leaching of chemical species, including Al, S, Cr, and Zn. However, this could potentially be explained by the lower volumes of water collected at the end of the tests (October). Calcium concentrations differed slightly among the cells, with average concentrations of about 379 mg/L, 440 mg/L, 431 mg/L, 470 mg/L, and 502 mg/L for C1, C2, C3, C4, and C5, respectively (Figure 2A). Average Mg concentrations were about 3756 mg/L, 1790 mg/L, 2151 mg/L, 39 mg/L, and 3.5 mg/L for C1, C2, C3, C4, and C5, respectively (Figure 2B). Average Mn concentrations were about 1421 mg/L, 45 mg/L, 44 mg/L, 1.5 mg/L, and 0.7 mg/L for C1, C2, C3, C4, and C5, respectively (Figure 2C). The average Al concentration was about 16 mg/L for C1 and did not exceed 0.45 mg/L for the amended cells (Figure 2D). The same trend was observed for Fe, which had an average concentration of 1034 mg/L in C1 and ≤ 0.2 mg/L for the amended tailings (Figure 2E). Average S concentrations were about 7523 mg/L, 2838 mg/L, 3289 mg/L, 707 mg/L, and 890 mg/L for C1, C2, C3, C4, and C5, respectively (Figure 2F). Zinc concentrations were increased in leachates from C1, with an average concentration of about 17.6 mg/L. Concentrations in the amended tailings did not exceed 1.4 mg/L (Figure 2G). Copper concentrations were higher for C1, with an average of 0.93 mg/L, while concentrations in the amended tailings were negligible (Figure 2H). Arsenic was only detected in leachates from C1, C2, and C3, with average concentrations of 0.28 mg/L, 0.29 mg/L, and 0.20 mg/L, respectively (Figure 2I). Finally, average Cr concentrations were about 0.42 mg/L, 0.028 mg/L, 0.029 mg/L, 0.001 mg/L, and 0.015 for C1, C2, C3, C4, and C5, respectively (Figure 2J).

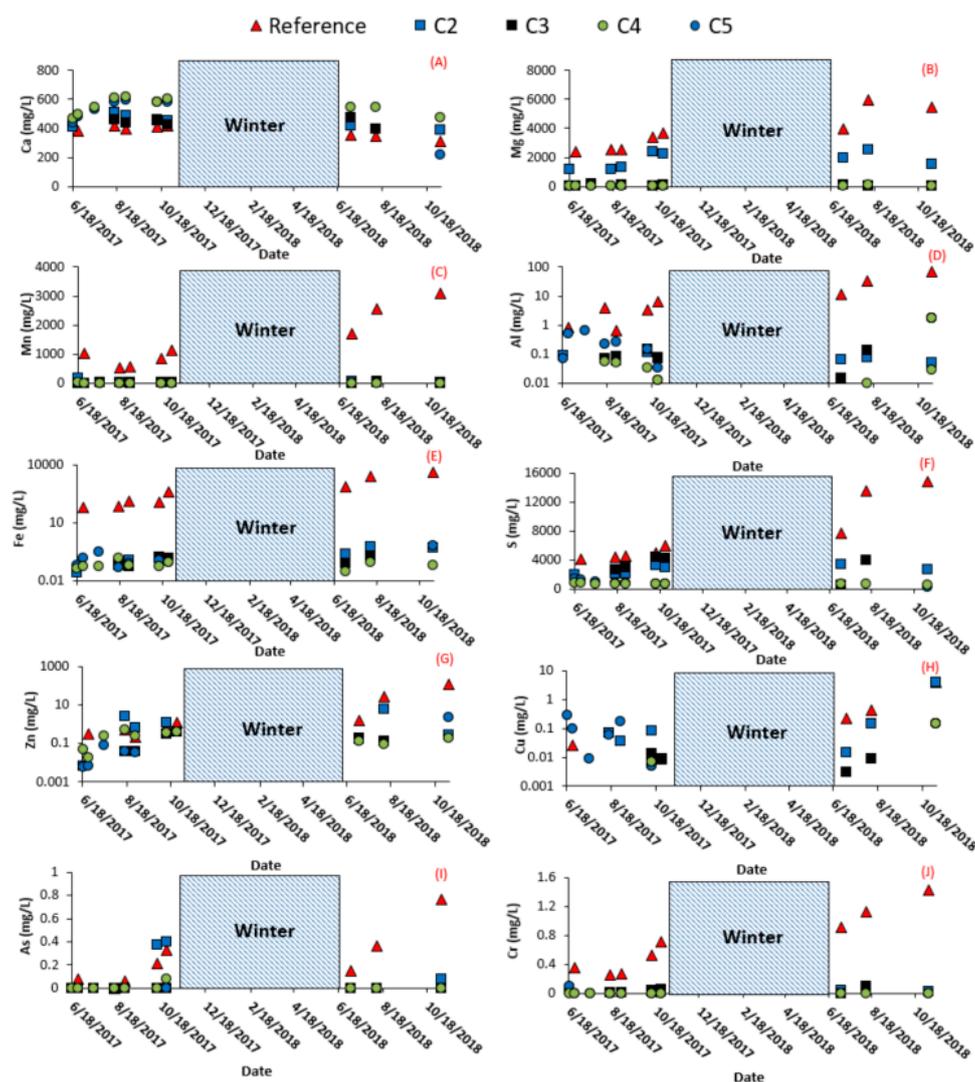


Figure 2. Evolution of concentrations of (A) Ca, (B) Mg, (C) Mn, (D) Al, (E) Fe, (F) S, (G) Zn, (H) Cu, (I) As, and (J) Cr in vertical infiltration waters.

3.2.2. Surface and Subsurface Runoff

Leachates collected from surface and subsurface runoff were analyzed for the same parameters as the leachates from vertical water infiltration. During the monitoring period, no leachates were collected from C4. Surface runoff from the control cell was acidic, with pH values less than 4 (Figure 3A), while the amended cells had pH values less than 6. Thus, the different amendment formulations successfully neutralized the acidity produced by the oxidized tailings. Average EC values for all cells were below 3.6 mS/cm (Figure 3B) and Eh values were higher than 100 mV, suggesting oxidizing conditions (Figure 3C).

Leachates from surface and subsurface runoff showed a different chemical quality compared to leachates from vertical infiltration. In general, the runoff leachates showed lower concentrations for most cations. Average Ca concentrations were about 357 mg/L, 403 mg/L, 444 mg/L, and 344 mg/L for C1, C2, C3, and C5, respectively (Figure 4A). Average Mg concentrations were about 66 mg/L, 153 mg/L, 79 mg/L, and 27 mg/L for C1, C2, C3, and C5, respectively (Figure 4B). Average Mn concentrations were about 27 mg/L for C1 and <13 mg/L for the amended tailings (Figure 4C). Average Al concentrations were <0.8 mg/L for all cells (Figure 4D). Iron was leached in relatively high concentrations from C1, with an average concentration of about 37 mg/L, but did not exceed 2 mg/L from the amended tailings

(Figure 4E). Average S concentrations were about 500 mg/L for all cells (Figure 4F). Zn, Cu, As, and Cr were detected in low concentrations (Figure 4G–J).

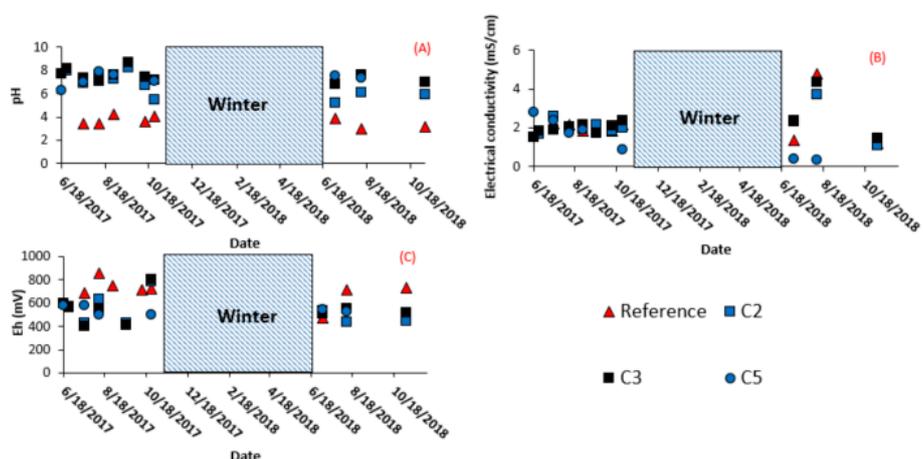


Figure 3. Evolution of (A) pH, (B) EC, and (C) Eh for leachates from surface runoff.

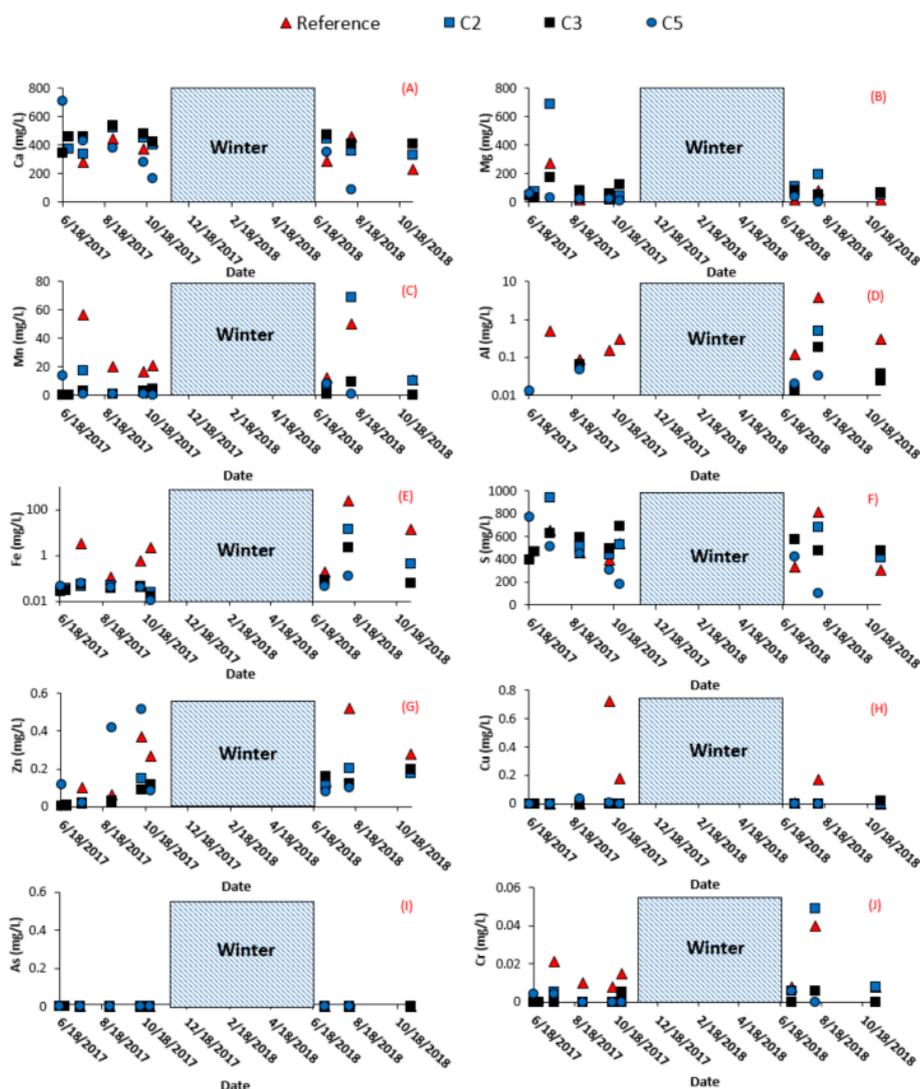


Figure 4. Evolution of concentrations of (A) Ca, (B) Mg, (C) Mn, (D) Al, (E) Fe, (F) S, (G) Zn, (H) Cu, (I) As, and (J) Cr for leachates from surface runoff.

3.3. Water Content Evolution

The results of water content monitoring for the five cells are illustrated in Figure 5. Water contents showed the same seasonal variation for all five cells. Water contents decreased during winter and increased during summer. Cell C5 (with 5% cement) showed the lowest VWC value at the end of the monitoring period. Cell C4 showed the highest VWC values over the two-year monitoring period. During the monitoring period, average VWCs were about 0.23, 0.40, 0.24, 0.43, and 0.26% for C1, C2, C3, C4, and C5, respectively. Additionally, the hydrogeological behavior of the cells observed in the field during sampling was very different. After a rain event, a water pool was observed within the cementitious cells. In addition, the volume of water collected in the barrel was always the smallest within C5 and C4, compared to that collected in the other cells.

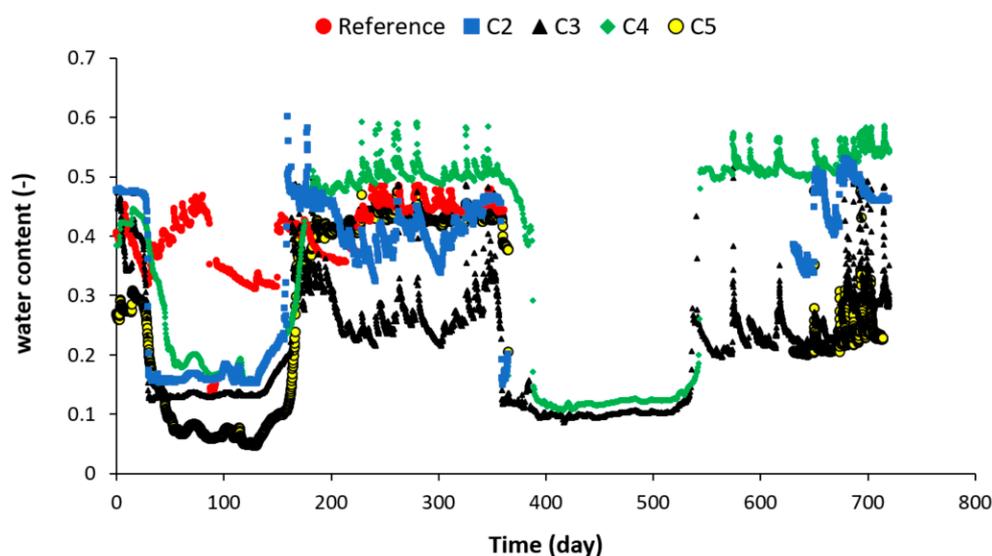


Figure 5. Evolution of volumetric water content (VWC) within the five field cells.

3.4. Field Cell Dismantlement

Samples taken from the field cells following the end of the monitoring period showed relatively similar mineralogical and chemical compositions (Table 3 and Figure 6A). The chemical compositions of the solid samples after the experiments were slightly variable among the cells. Al concentrations varied between 1.5 and 2.2 wt %, Fe concentrations were between 20 and 29 wt %, and S concentrations ranged between 4 and 7.5 wt % (other elements are presented in Table 3). The chemical differences among the samples are likely due to their different compositions (Table 1) and the initial differences due to the high volume of tailings used to fill the different cells.

Figure 6A shows a summary of the mineralogical compositions of the samples after two years of leaching tests. Sulfides consisted primarily of pyrite. Carbonates were mainly present as siderite, with moderate concentrations of calcite. Secondary minerals mainly included Fe oxides, gypsum, and Fe sulfates. The rest of the mineralogical composition comprised various silicates and other oxides.

Pyrite varied between 5 and 9 wt % and carbonate contents varied between 19 and 42 wt %. Calcite was detected in high concentrations in samples from the C3, which initially contained 10% limestone. Iron oxides were detected in high concentrations in samples from C1. This could be explained by precipitation of secondary Fe phases, due to the high Fe concentrations within the reference cell, especially as compared to the other cells (Figure 2A). Iron oxides were present in the form of coatings on the surface of carbonates, sulfides, and silicates. Gypsum contents varied between 0.5 and 6 wt %.

The mineralogical composition of the studied samples shows Ca occurring primarily in gypsum in the sample from C1 and in carbonates in samples from the amended cells (Figure 6B). More than 70 wt % of Fe occurred in siderite and Fe oxides in all samples (Figure 6C) and about 10 wt % of Fe occurred in pyrite. Sulfur was primarily associated with pyrite and gypsum (Figure 6D).

Sulfides and carbonates showed different degrees of liberation and associations (Figure 7). Mineral liberation analyses indicate the exposed areas of mineral. Pyrite was approximately 24, 13, 14, 28, and 18 wt % liberated in samples from C1, C2, C3, C4, and C5, respectively (Figure 7A). Pyrite liberation analyses for C4 and C5 only took into account pyrite locking by mineral phases and not pyrite locking by the cementation of the tailings. Therefore, the pyrite liberation within cementitious amendment was overestimated in this case. The majority of the pyrite consisted of binary associations with Ca–Mg carbonates and Fe oxides (Figure 8). Ca–Mg carbonates were about 0.62, 37, 51, 27, and 13 wt %, liberated in samples from C1, C2, C3, C4, and C5, respectively (Figure 7B). The majority of Ca–Mg carbonates consisted of binary associations with Fe oxides/siderite; they were about 24, 48, 44, 57, and 68 wt % for samples from C1, C2, C3, C4, and C5, respectively.

Table 3. Chemical composition of solid samples after field kinetic testing.

Elements (ppm)	Al	As	Ba	Be	Bi	Ca	Cd	Co	Cr	Cu	Fe	K	Li	Mg	Mn	Mo	Na	Ni	Pb	S	Ti	Zn
Detection limit (ppm)	60	5	5	5	5	60	5	5	5	10	10	1	5	15	5	5	1	5	5	200	25	55
C1	20,200	910	38	<5	108	37,230	39	<5	24	<10	272,400	2600	<5	1515	4237	8	10,700	25	128	75,350	627	99
C2	18,250	783	36	<5	108	50,910	37	<5	24	<10	256,300	2580	<5	7839	6104	7	9360	27	102	71,600	317	96
C3	15,360	654	31	<5	109	106,000	30	<5	22	<10	221,000	2530	<5	10,940	5122	9	7810	23	99	66,060	486	85
C3-duplicate	15,150	673	32	<5	105	100,900	30	<5	22	<10	208,900	2520	<5	11,100	5082	9	8090	24	98	64,600	501	83
C4	21,930	399	66	<5	104	33,660	37	<5	36	<10	291,200	2210	<5	9660	7698	6	12,900	21	105	57,920	865	115
C5	21,970	426	49	<5	100	40,720	38	<5	40	<10	281,800	2110	<5	7711	6775	8	12,600	21	136	43,100	756	107

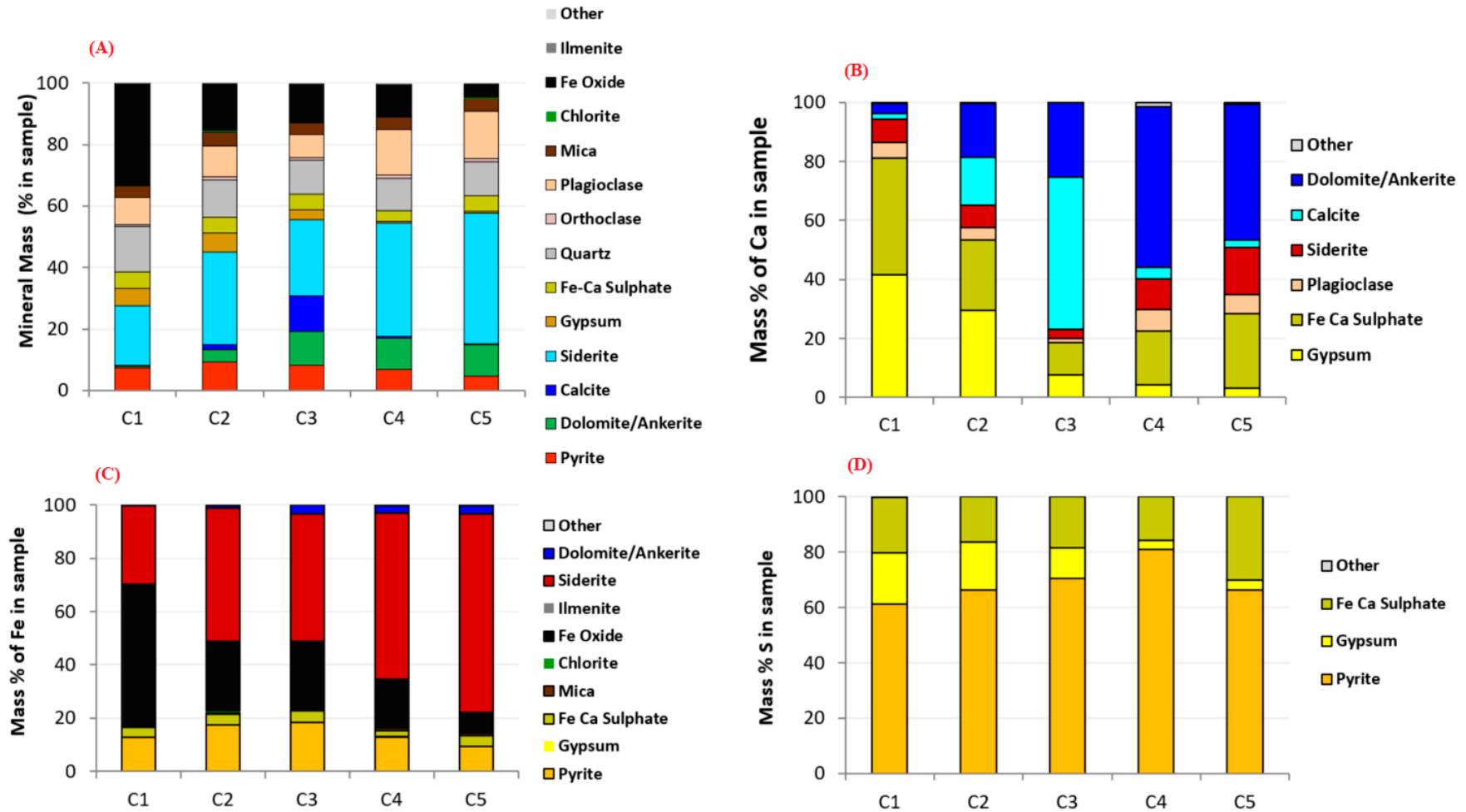


Figure 6. (A) Modal mineralogy for the dismantled samples; (B) Ca department; (C) Fe department; and (D) S department within the dismantled samples.

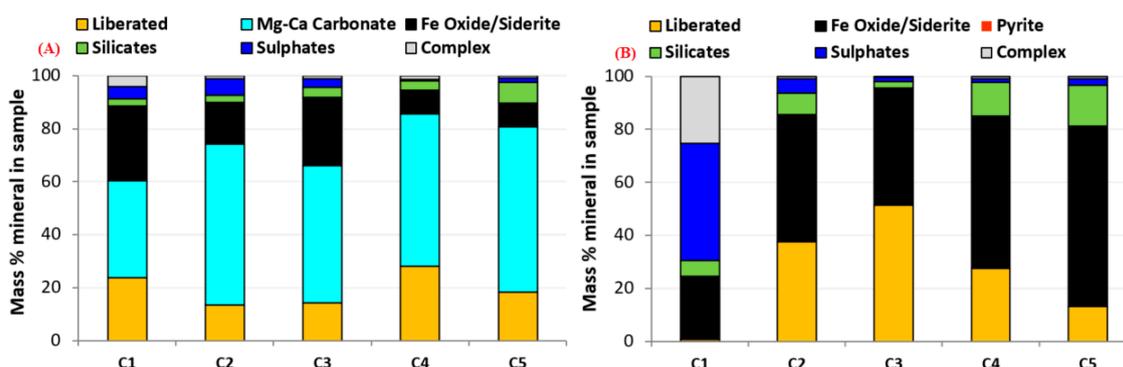


Figure 7. Degree of liberation for (A) carbonates and (B) sulfides.

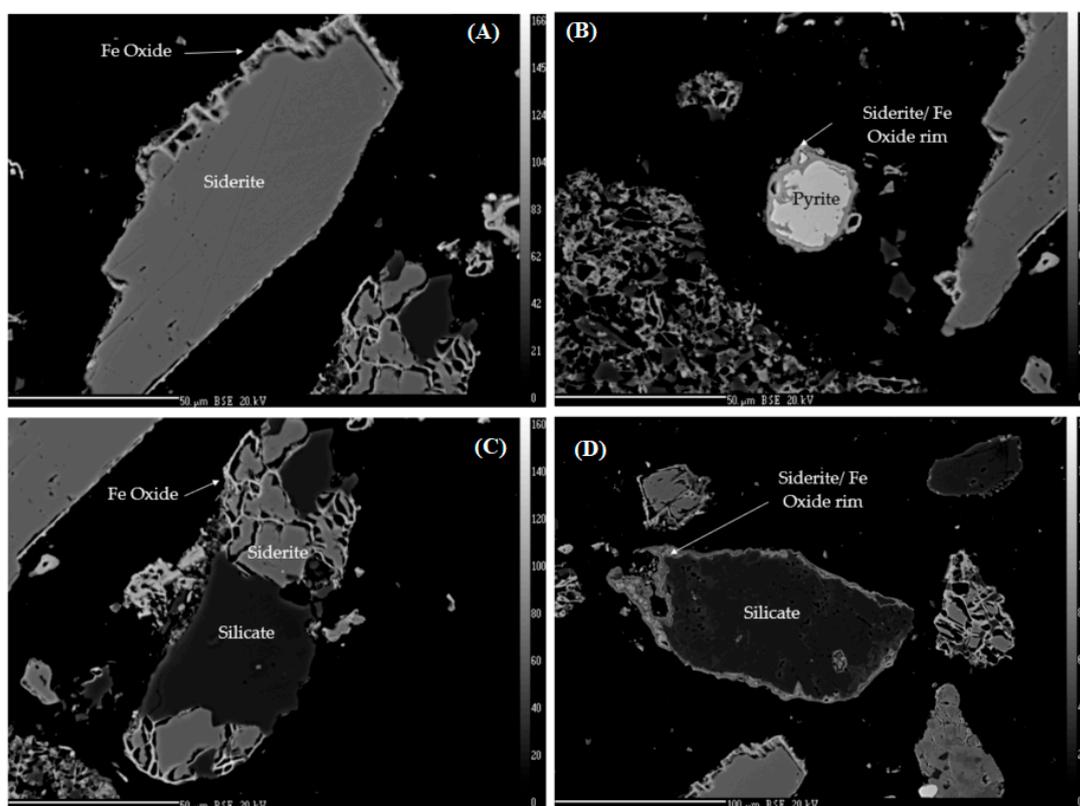


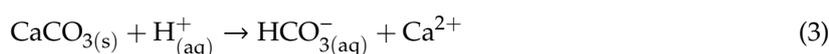
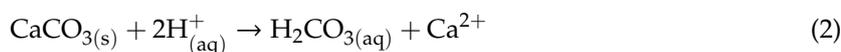
Figure 8. SEM images showing (A) siderite coated by Fe oxides, (B) pyrite coated by siderite and Fe oxides, and (C) and (D) silicates coated by Fe oxides and siderite.

4. Discussion

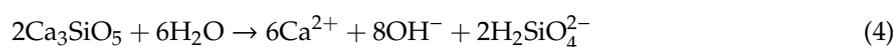
The oxidized tailings showed acidic pH values (<4) and, together with other monitoring data, the pH measurements suggest that surface area and contact time considerably influence final water quality. Overall, leachates from vertical infiltration samples were more acidic and more loaded in chemical species than those from runoff samples. This makes sense as the contact time between water and tailings is greater during infiltration than during surface and subsurface runoff. Vertical water infiltration rates depend upon the physical and hydrogeological properties of the tailings [62–66]. However, surface runoff rates are greatly controlled by the morphological properties of the TSF (e.g., slope gradient, slope length) and the precipitation intensity [67–69].

The geochemical behavior of amended tailings showed that alkaline and cementitious amendments are both promising techniques that could be used for stabilizing and neutralizing acid-generating

tailings. The leachates from the amended cells were characterized by circumneutral pH values and were less charged with chemical species than the control cell, except in the case of Ca, which was released more from cell C5 (5 wt % OPC) (Summary of chemical analysis is presented in Table S1). In fact, the dissolution of carbonates and OPC produces alkalinity and buffers the pH of leachates [11,15,31,38,57–59,70–79]. The pH was established early, due to high dissolution rates of limestone and cement in the acidic media. The dissolution of carbonates (calcite) is explained by Equations (2) and (3):



Ca leaching was higher within C5, which contained OPC as a binder. In fact, the dissolution of tricalcium silicate (C3S), which is contained in OPC, is explained with the following reaction (Equation (4)):



Otherwise, Ca was released by gypsum and by carbonate dissolution in the control cell.

Other chemical species (As, Fe, Al, Li, Pb, Cr, Li, and Zn) were immobilized within the amended cells. The mechanisms responsible for the attenuation of these metal(-oid)s include: Precipitation, co-precipitation, and sorption related to the formation of secondary Fe oxyhydroxides at circumneutral pH values [2,11,24]. The Fe–S Eh–pH diagram illustrated in Figure 9 shows that Fe should precipitate as Fe oxides under the conditions observed in this study. However, for the cementitious amendments (C4 and C5), in addition to the precipitation of secondary Fe phases, two other mechanisms could be responsible for the attenuation of chemical species: (i) Physical trapping and (ii) reduction of water/tailings contact surface, due to hardening of the tailings by binders (FA and OPC). The application of cementitious additives enhances the mechanical resistance of tailings and increases their long-term impermeability [20,44,60,61]. As shown in Figure 10, all amendments used in this study were capable of decreasing the mobility of the analyzed metal(-oid)s, compared to the reference cell (with the exception of As and Cr in the vertical infiltration and runoff leachates from C2, respectively). A reduction factor was calculated for each analyte using Equation (5):

$$RF = 100 \times \left(1 - \frac{C_a}{C_0}\right) \quad (5)$$

where RF is the reduction factor, C_a is the concentration of analyte in the amended cell, and C_0 is the concentration of the analyte in the reference cell.

The amendments reduced the S release from the C2, C3, C4, and C5 with respect to the control. However, more S was released from C2 and C3 than from C4 and C5. This could potentially be explained by the hardening of the tailings–cement mixture. Limestone amendments reduce the leaching of chemical species only by chemical process (e.g., precipitation of secondary phases); however, cementitious amendments reduce the leaching of chemical species by both chemical and physical processes (e.g., reducing the surface contact).

Iron was almost completely immobilized within the different amendment formulations, with RF values between 80 and 100%. Aluminium concentrations were reduced by more than 90%, with a maximum RF of 100% for C3, C4, and C5. Fe oxides precipitated in all cells, as illustrated in the mineralogical composition presented in Figure 6A. However, in C1, there were higher iron oxide concentrations. This could be explained by the pH–Eh conditions. The pH in the reference cell was acidic, which means high reactivity of sulfides would be expected under these conditions. Zinc mobility was greatly reduced for infiltrating waters, however surface and subsurface runoff waters showed RF values between 23 and 67%. Arsenic concentrations were completely reduced for the cementitious formulations with an RF of 100%, except for the formulation C3 (10% limestone) where it still released. Lead was successfully immobilized within RF values greater than 77% in all cases. The cells with

cement showed the highest effectiveness regarding the As immobilization, due to: Hardening of tailings, which reduces the reactive surface, when the As is trapped within stable C–S–H phases [80]. The cementitious amendments offer more mechanisms to reduce the As leaching compared to alkaline amendments. Nickel showed RF values greater than 53%. Finally, Cr was successfully stabilized by amendments C3, C4, and C5, while C2 showed appreciable concentrations in surface and subsurface runoff samples. The overall order of effectiveness of the different amendments was $C5 = C4 \geq C3 \geq C2$. Therefore, the cementitious amendments appear to show a greater capacity to immobilize chemical species with respect to the tested alkaline amendments. The incorporation of fly ash to partially substitute the cement could reduce the costs related to the use of cementitious amendments. Considering both economic factors and overall effectiveness, the formulation with fly ash could be considered the best option. Furthermore, the cementitious amendments include the physical trapping of contaminants such as arsenic [44,47]. Indeed, the evolution of the VWC presented in Figure 5 shows that cell C5 had the lowest water content at the end of the test. This suggests that the infiltration speed is lower due to the hardening of the tailings.

Quantitative mineralogical analyses of the five cells showed that carbonates were more liberated compared to sulfides (Figure 7). Mineral liberation is recognized as a key factor in the reactivity of mine wastes [81–83]. After two years of monitoring, the reactive fraction of carbonates was higher than that of the sulfides.

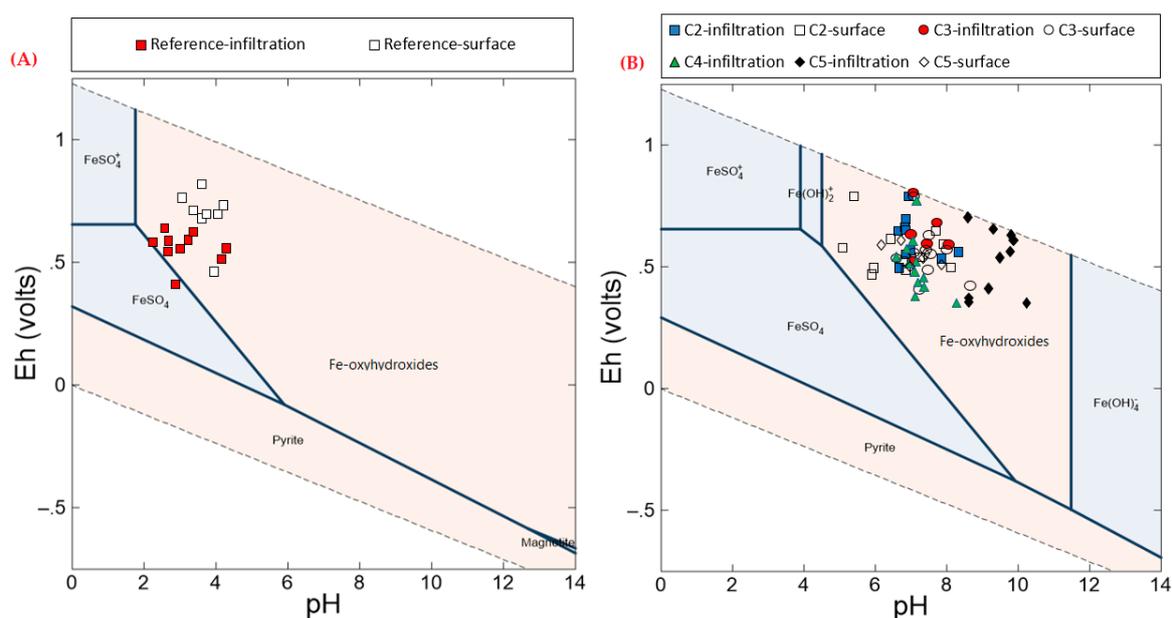


Figure 9. Fe and S pH–Eh diagrams for the leachates from (A) the reference cell and (B) amended tailings.

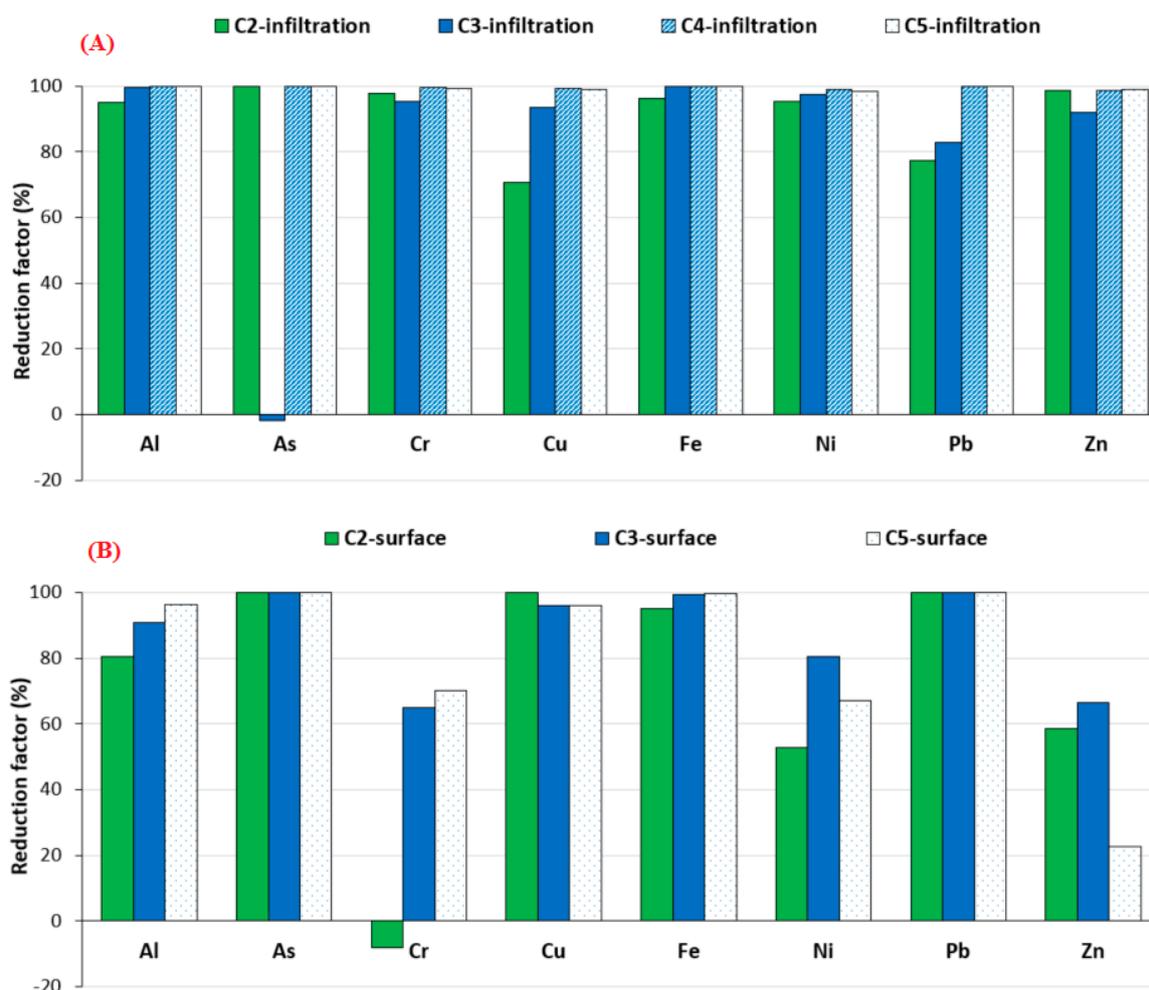


Figure 10. Reduction factors for Al, As, Cr, Cu, Fe, Ni, Pb, and Zn for the different amendment formulations. (A) Vertical infiltration leachates; (B) surface and subsurface runoff leachates.

5. Conclusions

Reducing risks, such as acid generation and contaminant release from mine tailings, is a serious challenge facing the mining industry and requires the development of cost-effective techniques and technologies. The application of limestone as an alkaline amendment or ordinary Portland cement and fly ash as cementitious amendments showed promising results to neutralize the acid produced and contaminants released from oxidized tailings, based on short-term, in situ tests. Furthermore, OPC was successfully partially substituted with fly ash, which could reduce the costs associated with cementitious amendments. In this study, cementitious amendments showed a high effectiveness compared to alkaline amendments. In fact, metal and metalloid concentrations were reduced more by both tested cementitious amendments than by the tested limestone amendments, which failed to attenuate Zn and As releases when applied at 5% and 10%, respectively. However, the applicability of mining amendments as a reclamation technique must consider two major points: i) The long-term behavior of amended tailings and for how long the contaminants will be immobilized, and ii) the cost related to applying the amendment. Economic factors will be greatly affected by the availability of the amendment materials at a close proximity to the mine site.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2075-163X/9/5/314/s1>, Figure S1: Images showing field cells construction and amendments mixing with oxidized tailing, Figure S2: Schematic representation of the cells constructed in the TSF (image not to scale), Figure S3: Calibration curves of the volumetric water content probes, Figure S4: Average daily temperature, Table S1: Summary of chemical

concentrations of the leachates from the different cells (\leq LD means that the concentration was inferior to the detection limit of ICP-AES).

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